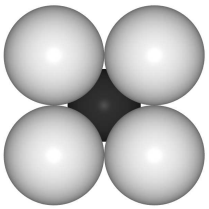
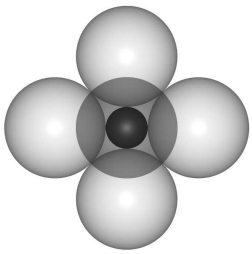
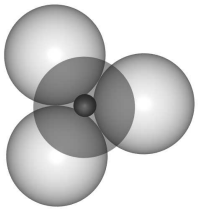
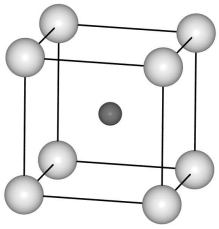
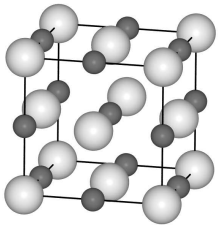
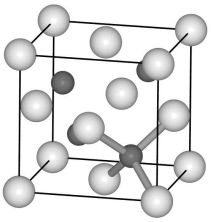
2. Alkalimetalle (Forts.)

2.3. Halogenide

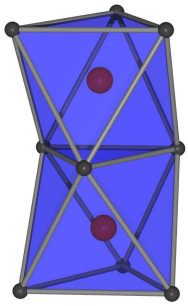
$CN_{Anion/Kation}$=	8	6	4
$\frac{r_{Kation}}{r_{Anion}}$	> 0.73	0.73 – 0.41	0.41 –0.22
Kationen-koordination			
Elementarzelle			
Strukturtyp	**CsCl**	**NaCl**	**ZnS** (Zinkblende)
M^{I}-Halogenide	CsCl, CsBr, CsI	LiF, LiCl, NaF, NaCl, KF, KCl, RbF, RbCl, CsF	-
M^{II}-Chalkogenide	-	MgO, CaO, SrO, BaO, CaS, SrS	BeO, MgTe

Ionenkristalle (AB-Typen): Koordinationszahlen, Strukturen, Verbindungen

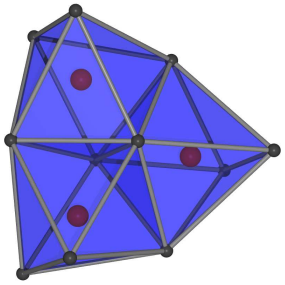
2.4. Oxide

Verb.	Name	Anion	$r_{O\text{-}O}$	Beispiel	Darstellung
$M^{I}_2O_2$	Peroxide	O_2^{2-}	149	Na_2O_2	Na (K): Verbrennen an Luft
$M^{I}O_2$	Hyperoxide	O_2^{-}	133	KO_2	K, Rb, Cs: Verbrennen an Luft
M_4O_6	Sesquioxide	$(O_2^{-})_2(O_2^{2-})$			(schwarz)
$M^{I}O_3$	Ozonide	O_3^{-}	135	RbO_3	Darst. nur mit O_3 (orange)

Sauerstoffverbindungen der Alkalimetalle, mit O−O-Bindungen

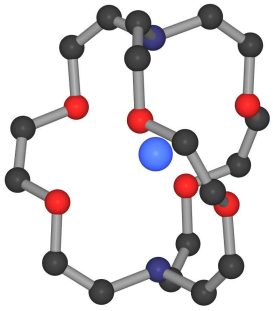


Rb_9O_2

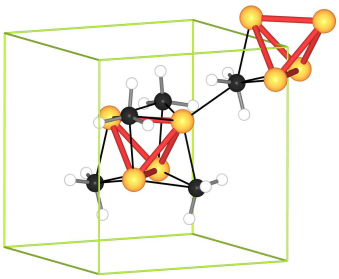


$Cs_{11}O_3$

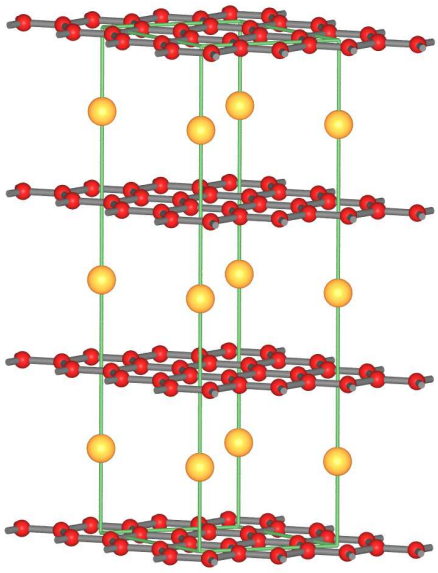
Cluster als Bauelemente der Suboxide



2.7. $[Na(222-crypt)]^+$



Struktur von *Me*Li



Struktur von CsC_8